

S.#	Question	Answer Option (A)	Answer Option (B)	Answer Option (C)	Answer Option (D)
1	How many times Mg atom is heavier than Carbon atom?	2 times	3 times	12 times	4 times
2	The reagent which leaves in the reaction mixture after reaction has completed is called as ____?	Limiting reactant	Excessive reagent	Consumed reagent	All of these
3	When we calculate radii of Hydrogen atom by this equation $r = .529A(n^2)$ where $n = 1, 2, 3, 4$, the distance between orbits of hydrogen atom will	decrease	increase	remains same	be constant
4	Spectrum is the visual display or _____ of component of white light when it is passed through prism	rarefaction	radiation	collection	dispersion
5	Which of the following shows Planck's quantum theory	$E=h\nu$	$E=c/v$	$E = cv$	none of these
6	A gas having volume of 1 dm ³ is enclosed in a vessel at 1 c and 2.5 atm. This gas is allowed to expand until new pressure is 2 (No Suggestions) will be new volume if the temperature is maintained at 273 k?	12dm ³	12.5dm ³	1 dm ³	12 .3dm ³
7	If a graph is plotted between pressure on x- axis and volume on y- axis for Boyle's law verification, the curve obtained is called	pseudoderm	isotherm	biotherm	all of these
8	The properties of gases, liquids and solids can be understand by	atomic theory	potential molecular theory	kinetic molecular theory	none of these
9	The forces present between the molecules are called as ?	Chemical bond	Intermolecular forces	Intra molecular forces	Strong forces
10	Which force present in a mixture of H ₂ and HCl molecules?	Dipole Dipole forces	Debye forces	London Dispersion forces	None of these
11	Which instrument is used to measure the vapour pressure of a liquid?	Barometer	Manometer	Thermometer	Sphygmometer
12	In polar solvents ionic crystals react	Slowly	fast	reacts on heating	No reaction
13	Melting point of crystalline solids are	Fixed	Variable	have a range of value	None of these
14	Lattice Points are also called	space lattice	crystal lattice	lattice sites	Lattice location
15	In an Irreversible reaction the tendency of it to go in reverse direction is	High	low	negligible	none Of these
16	Equilibrium constant has	Units	No Units	Both A and B	A negative value
17	The equilibrium Constant is always written as a ratio of	Reactants over products	Products over reactants	Product times Reactants	None of these
18	Example of fast reactions are	Neutralization	Double Decomposition Reactions	Ionic Reactions	All of these
19	Activation energy for forward and backward reactions is always	Same	Different	Reciprocal of each other	Higher for forward reaction
20	Following are slow reactions except?	Rusting of iron	Weathering of stones	Reaction between silver nitrate and sodium chloride	Fermentation of sugar
21	An enthalpy cycle used to calculate the lattice energy is	Hess Law	Carnot Cycle	Born haber Cycle	Haber's Process
22	Hess law is known as	Law of heat summation	Law of constant heat capacity	Law of mass action	Law of conservation of mass
23	Work done by the system is always	Positive	Zero	Negative	Equals to Unity
24	Standard Hydrogen Electrode is dipped into	1 molar solution of NaOH	1 molar solution of HCl	Nitric Acid	KOH
25	Standard Reduction potential is found by dipping the electrode in its own ion with molarity equals to	0.1	1	1.5	0
26	In Industry Caustic Soda is formed by electrolysis of	Dilute NaCl Solution	Fused NaCl	Concentrated NaCl Solution	NaCl Solution
27	The covalent radius of Hydrogen is	37.6pm	37.7pm	37.7nm	37.6nm
28	The minimum amount of energy required to remove an electron from its gaseous atom to form an ion is known as	electron affinity	ionization energy	electronegativity	potential energy
29	The sum of cationic and anionic radius in a crystal lattice is equal to	inter-titan distance R	inter ionic distance R	inter- cationic distance R+	inter-anionic distance R-
30	Which of the following is NOT an alkali metal	Li	Fr	H	Cs
31	Which oxidation state is shown by alkali metals	-1	-2	1	2
32	Which elements form alloys	Alkali metals	Alkaline earth metals	Transitions metals	Halogens
33	Transition elements have properties of _____?	Metals	None metals	Between both metals and nonmetals	None of these
34	The species which donates electrons to the central metal atom in the coordination sphere is called	Anion	Cation	The ligand is positively charged	Acid
35	Compounds with same molecular formula but different structural formula and properties are called:	Isomers	Organic compounds	Monomers	Tautomers
36	What is the origin of organic compounds?	Plants	Animals	Anything except living organisms	Both a) and B)
37	Isomerism becomes possible when:	There more than two carbon atoms	There more than three carbon atoms	There more than three hydrogen atoms	There more than three in the atoms in the molecule
38	the catalytic oxidation of alkane is used	as fuel	in organic synthesis	prepare higher fatty acids	all of these
39	The stability of benzene is due to	cyclic structure	that it gives addition reaction	delocalization of pi electrons	none
40	Which of the following belongs to the class of saturated hydrocarbons	Propene	Butane	Isobutylene	Ethyne
41	Victor Grignard was awarded a Nobel Prize in	Physics	Chemistry	Mathematics	Quantum chemistry
42	In primary alkyl halide, halogen atom is attached to a carbon atom which is	Attached to one or no carbon	Attached to two or no carbon	Attached to one or two carbons	Attached to one or three carbons
43	Chloroethane is an example of	Primary alkyl halide	Secondary alkyl halide	Tertiary alkyl halide	None of the above
44	In cannizzaro's reaction, Benzyl alcohol is obtained from	benzaldehyde	benzene	picric acid	benzyl chloride
45	The compounds which are formed by the replacement of one of the H of water by a alkyl group are called as ____?	Ethers	Phenols	Alcohols	Carboxylic acids
46	The alcohols in which the carbon which is attach to the OH group is further attach with two carbon atoms is called as?	Primary alcohols	Secondary alcohols	Tertiary alcohols	None of these
47	Bulky ketones do not react with	Sodium bisulphite	HCl	Grignard reagent	HCN
48	If both alkyl groups attached to carbonyl in ketone are same then it is called as ____?	Mixed	Unsymmetrical	Symmetrical	None of these
49	In IUPAC nomenclature, aldehydes are named as _____?	Alkanol	Alkanal	Alkanone	Alkaldehyde
50	Which of the following is correct about IUPAC naming of carboxylic acid?	In the parent name COOH position is not mentioned	COOH is written as substituent and position is mentioned	COOH is given 1st position in all carboxylic acid	Both a and c
51	Carboxylic acids are produced by the oxidation of _____?	Alcohol	Aldehyde	Ketone	All of these
52	The compounds which have -CN group are called as ____?	Aldehydes	Imines	Amines	Nitriles
53	Amino acids in proteins are held together by:	Peptide bond	Phosphodiester linkage	H-bonding	Ether linkage
54	What is the Optimum temperature for working of enzyme in human body?	40°C	35°C	37°C	32°C
55	After digestion proteins change into:	Amino acids	Starch	Glycogen	Lipids
56	The name protein is derived from _____ word proteios meaning_____.	Greek, Important.	French, Prime Importance	Greek, Functional	Greek, Prime Importance